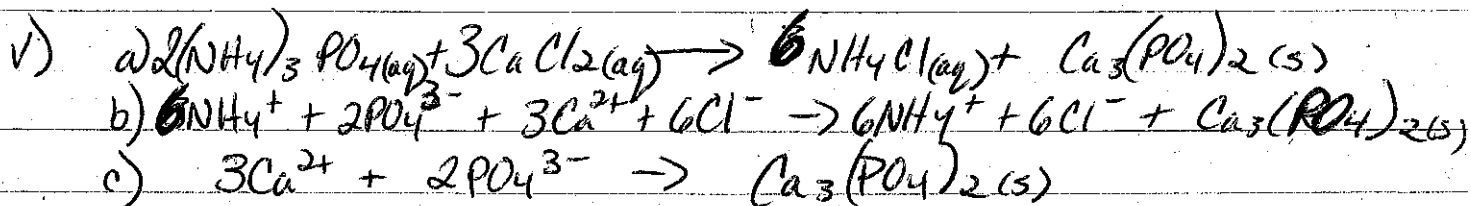
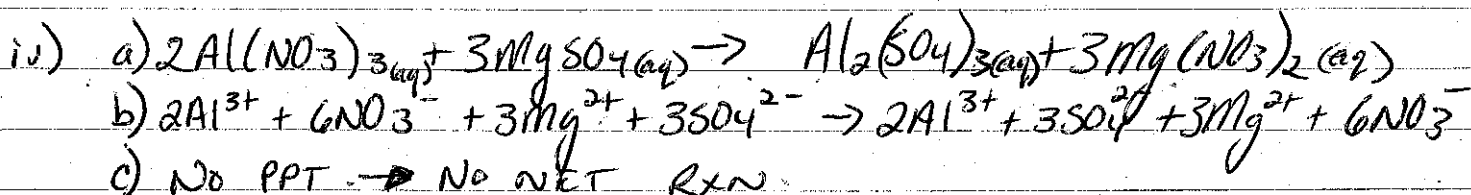
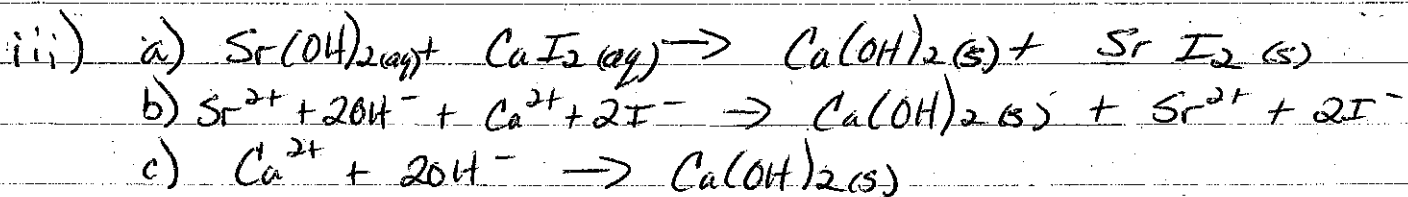
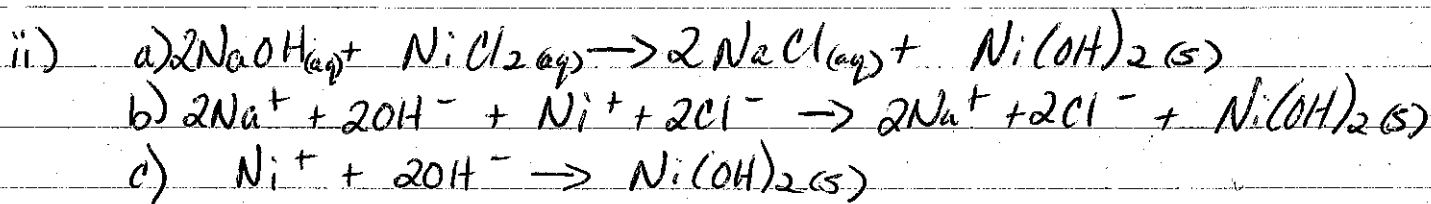
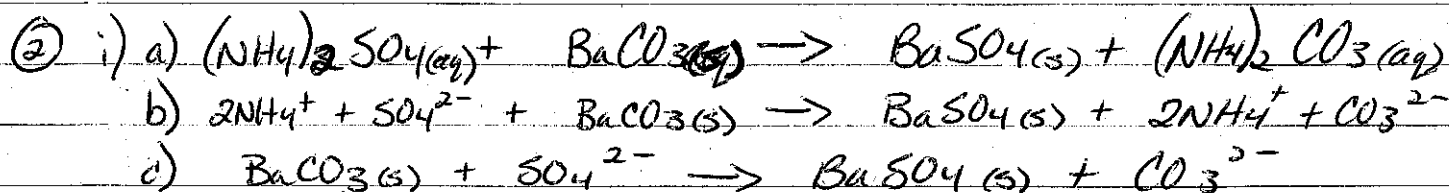


# AQUEOUS RXNS REVIEW KEY

- ① a) ~~sol.~~ ~~insoluble~~ d) Soluble  
 b) Soluble e) Soluble  
 c) Low sol.  $Mg_3(PO_4)_2$  f) Soluble



③ See NOTES!

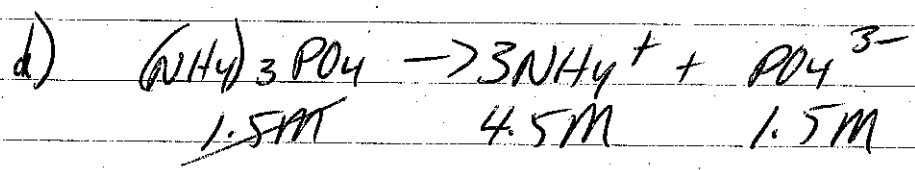
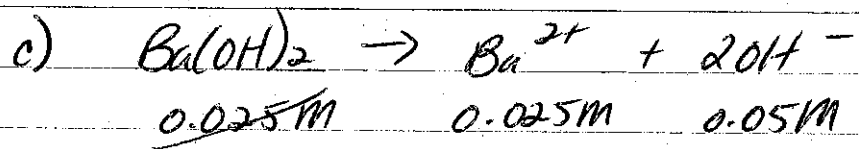
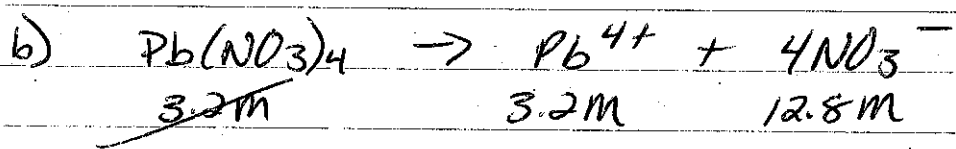
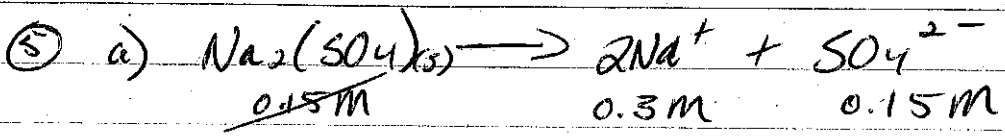
IONIC  $\rightarrow$  dissociate (break apart)  
 Covalent  $\rightarrow$  surrounded by  $H_2O$  as a whole molecule

④ a) weak electrolyte → substance that dissolves, but only dissociates a little bit.

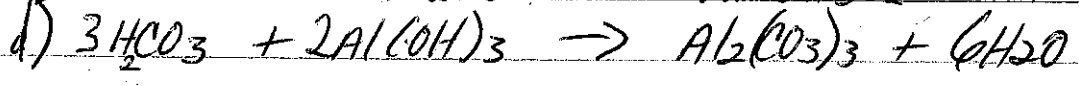
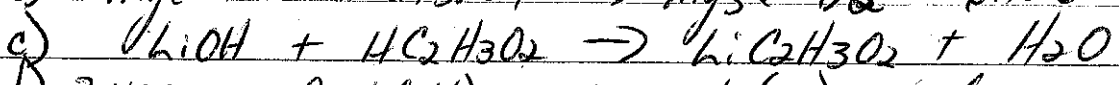
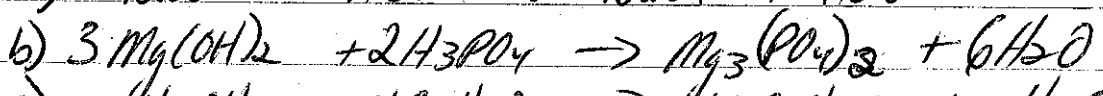
ex) Vinegar.

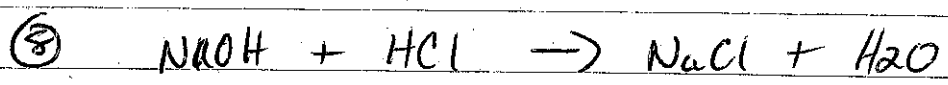
b) Strong electrolyte → substance that dissociated completely when dissolved ex) HCl

c) Non electrolyte → substance that doesn't dissociate when dissolved ex) C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>

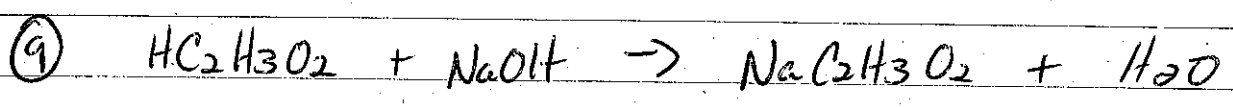


⑥ See notes.

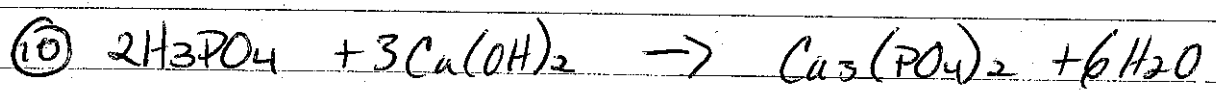




0.5 M x 0.03 L = 0.015 mol HCl x 1/1 = 0.015 mol NaOH = 0.05 L = 0.3 M



0.48 M x 0.0525 L = 0.0252 mol NaOH x 1/1 = 0.0252 mol HC2H3O2 = 0.017 L = 1.48 mol/L

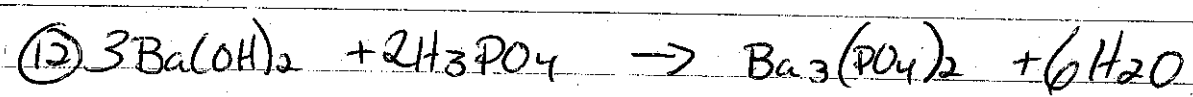


0.0053 mol/L x 0.022 L = 1.17 x 10^-4 mol Ca(OH)2 x 2/3 = 7.77 x 10^-5 mol H3PO4

7.77 x 10^-5 mol / 0.0025 mol/L = 0.0311 L

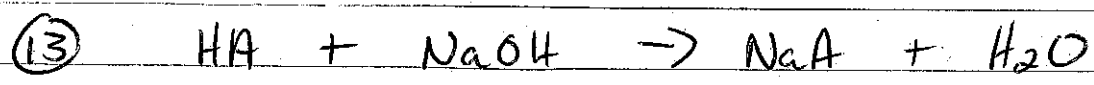


2.45 g NaOH / 40.09 g/mol = 0.061 mol NaOH x 1/2 = 0.0306 mol H2SO4 = 2.2 mol/L = 0.014 L



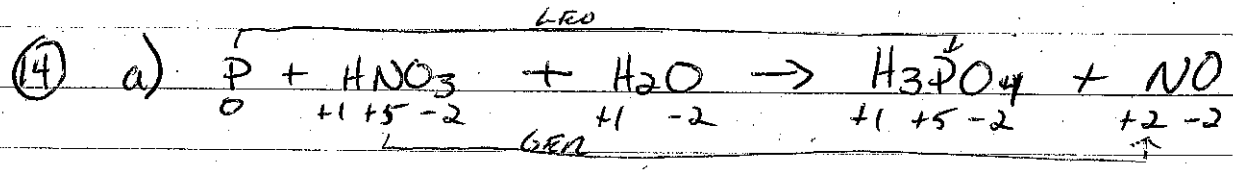
0.25 mol/L x 0.075 L = 0.01875 mol H3PO4 x 3/2 = 0.0281 mol Ba(OH)2

0.0281 mol Ba(OH)2 x 171.32 g/mol = 48.14 g

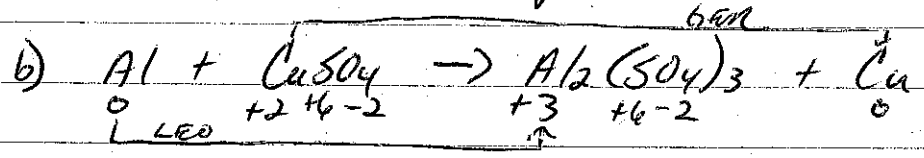


$0.175 \frac{\text{mol}}{\text{L}} \times 0.025 \text{L} = 0.004375 \text{mol NaOH} \times \frac{1}{1} = 0.004375 \text{mol AA}$

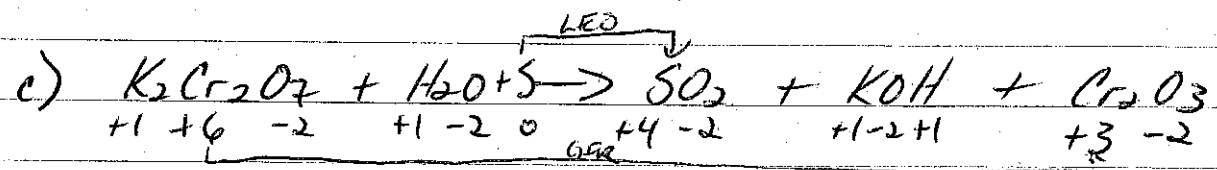
$\frac{0.900 \text{g}}{0.004375 \text{mol}} = \boxed{205.71 \text{g/mol}}$



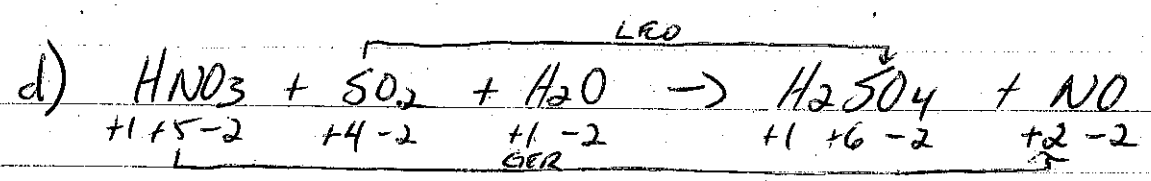
Reducing Agent = P  
 Oxidizing Agent = HNO<sub>3</sub>



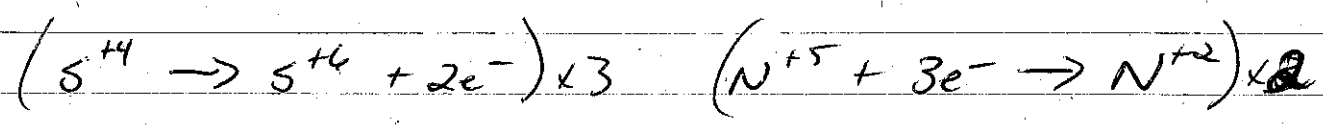
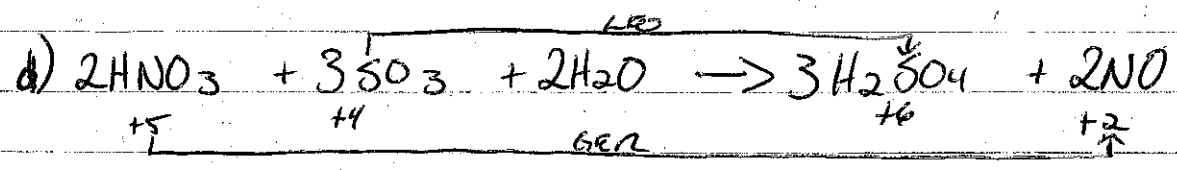
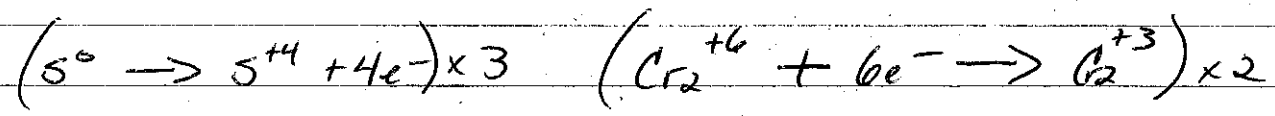
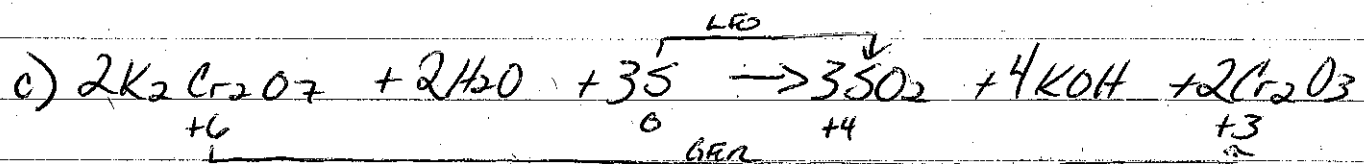
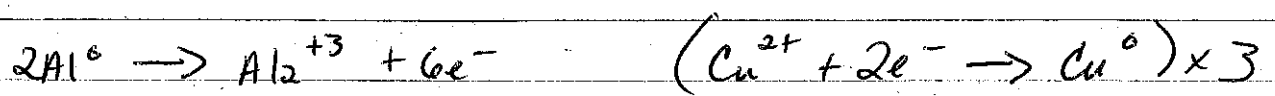
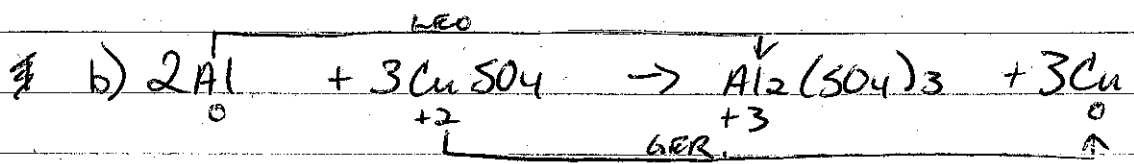
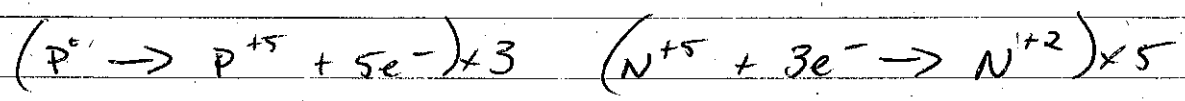
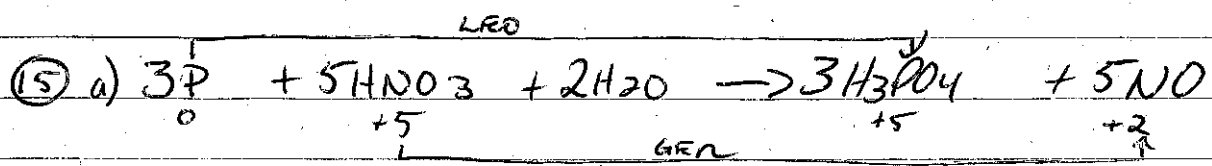
R.A = Al  
 O.A = CuSO<sub>4</sub>



R.A = K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>  
 O.A = S



R.A = SO<sub>2</sub>  
 O.A = HNO<sub>3</sub>



~~5~~

